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[Edvantage Science Ap Chemistry 2 Answer Key](#)

1. NOCl decomposes to form NO and Cl₂ by second-order reaction kinetics with $k = 0.020 \text{ L mol}^{-1} \text{ s}^{-1}$. How long will it take for a sample of NOCl with an initial concentration of 0.038 M to drop to 0.012 M? (From Edvantage Science AP Chemistry 2)



$$\frac{1}{[\text{NOCl}]_t} - \frac{1}{[\text{NOCl}]_0} = kt$$

$$\frac{1}{0.012} - \frac{1}{0.038} = (0.020)t$$

$$t = 2850.98 = 48 \text{ min}$$

2. Pharmacokinetics is the study of how the body takes up, distributes, and eliminates drugs. The elimination of a newly designed Canadian drug from the body follows second-order kinetics. The drug is called Zippidy Eh. Its molar mass is 122.0 g/mol. The rate constant for the elimination process is $2.4 \text{ L mol}^{-1} \text{ min}^{-1}$. If the initial dose of Zippidy Eh is 250 mg for a 70 kg patient with a blood volume of 4.8 L, what mass of the drug will remain 12 h after ingestion? (From Edvantage Science AP Chemistry 2)

$$0.25 \text{ g} \times \frac{1 \text{ mol}}{122.0 \text{ g}} \times \frac{1}{4.8 \text{ L}} = 0.0004261 \text{ M}$$

$$0.00024567 \text{ mol} \times 4.8 \text{ L} \times \frac{122.0 \text{ g}}{1 \text{ mol}}$$

$$\frac{1}{[Z]_t} - \frac{1}{[Z]_0} = kt$$

$$= 0.14 \text{ g}$$

$$\frac{1}{[Z]_t} - \frac{1}{0.0004261} = (2.4 \text{ L mol}^{-1} \text{ min}^{-1})(720 \text{ min})$$

$$[Z]_{720} = 0.00024567 \text{ M}$$

check that this makes sense

3. Dinitrogen pentoxide is the unstable anhydride of nitric acid that decomposes by first-order kinetics into nitrogen dioxide and oxygen gas. The rate constant is $6.22 \times 10^{-4} \text{ s}^{-1}$ at 45°C. If the initial concentration of nitrogen pentoxide is 0.250 M, what is the concentration after exactly 1 hour? (From Edvantage Science AP Chemistry 2)



$$\ln[\text{N}_2\text{O}_5]_t - \ln[\text{N}_2\text{O}_5]_0 = -kt$$

$$\ln[\text{N}_2\text{O}_5]_t - \ln 0.250 = -(6.22 \times 10^{-4} \text{ s}^{-1})(3600 \text{ s})$$

$$\ln[\text{N}_2\text{O}_5]_t = -3.6255$$

$$[\text{N}_2\text{O}_5] = e^{-3.6255} = 0.0266 \text{ M}$$



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